

**Bonding Practice**

**STEP 1:** For each compound, **decide** whether it is ionically bonded or covalently bonded.

 **Label** the bond type. How do I do this?

**STEP 2:** **Draw** the Lewis dot structure for each element.

 How do I do this?

**STEP 3:** Ionic? **Show** movement of electrons AND label the charge on each atom.

Covalent? **Make** a second drawing **showing** electrons in the bonded molecule.

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| K2O  | SrF2  |
| CaBr2  | I2  |
| CH4­  | MgF2  |
| HBr  | FeO (Fe = 2 valence electrons)  |
| BaO  | Challenge: C2H6  |