

**Bonding Practice**

**STEP 1:** For each compound, **decide** whether it is ionically bonded or covalently bonded.

**Label** the bond type. How do I do this?

**STEP 2:** **Draw** the Lewis dot structure for each element.

How do I do this?

**STEP 3:** Ionic? **Show** movement of electrons AND label the charge on each atom.

Covalent? **Make** a second drawing **showing** electrons in the bonded molecule.

|  |  |
| --- | --- |
| K2O | SrF2 |
| CaBr2 | I2 |
| CH4­ | MgF2 |
| HBr | FeO  (Fe = 2 valence electrons) |
| BaO | Challenge: C2H6 |